

# Die elektrochemische Spannungsreihe

| Element im Redoxpaar, dessen Oxidationsstufe sich ändert | oxidierte Form + z e <sup>-</sup> ⇌ reduzierte Form | Standardpotential E <sup>0</sup> |
|--|---|----------------------------------|
| Fluor  | $F_2 + 2e^- \rightleftharpoons 2F^-$                | +2,89V                           |
| Chlor  | $Cl_2 + 2e^- \rightleftharpoons 2Cl^-$              | +1,396V                          |
| Brom   | $Br_2 + 2e^- \rightleftharpoons 2Br^-$              | +1,098V                          |
| Silber   | $Ag^+ + e^- \rightleftharpoons Ag$                  | +0,799V                          |
| Iod  | $I_2 + 2e^- \rightleftharpoons 2I^-$                | +0,535V                          |
| Kupfer(I)  | $Cu^+ + e^- \rightleftharpoons Cu$                  | +0,518V                          |
| Kupfer(II)   | $Cu^{2+} + 2e^- \rightleftharpoons Cu$              | +0,339V                          |
| Wasserstoff  | $2H_3O^+ + 2e^- \rightleftharpoons 2H_2O + H_2$     | 0V                               |
| Eisen(III)   | $Fe^{3+} + 3e^- \rightleftharpoons Fe$              | -0,037V                          |
| Blei   | $Pb^{2+} + 2e^- \rightleftharpoons Pb$              | -0,126V                          |
| Zinn   | $Sn^{2+} + 2e^- \rightleftharpoons Sn$              | -0,141V                          |
| Chrom  | $Cr^{3+} + 3e^- \rightleftharpoons Cr$              | -0,89V                           |
| Zink   | $Zn^{2+} + 2e^- \rightleftharpoons Zn$              | -0,762V                          |
| Aluminium  | $Al^{3+} + 3e^- \rightleftharpoons Al$              | -1,677V                          |
| Natrium  | $Na^+ + e^- \rightleftharpoons Na$                  | -2,714V                          |
| Calcium  | $Ca^{2+} + 2e^- \rightleftharpoons Ca$              | -2,868V                          |
| Lithium  | $Li^+ + e^- \rightleftharpoons Li$                  | -3,040V                          |

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