Die elektrochemische Spannungsreihe

Element im Redoxpaar, dessen Oxidationsstufe sich ändert	**oxidierte Form + z e^- ⇌ reduzierte Form**	**Standardpotential E^0**
Fluor	$F_2 + 2e^- \rightleftharpoons 2F^-$	+2,89V
Chlor	$Cl_2 + 2e^- \rightleftharpoons 2Cl^-$	+1,396V
Brom	$Br_2 + 2e^- \rightleftharpoons 2Br^-$	+1,098V
Silber	$Ag^+ + e^- \rightleftharpoons Ag$	+0,799V
Iod	$I_2 + 2e^- \rightleftharpoons 2I^-$	+0,535V
Kupfer(I)	$Cu^+ + e^- \rightleftharpoons Cu$	+0,518V
Kupfer(II)	$Cu^{2+} + 2e^- \rightleftharpoons Cu$	+0,339V
Wasserstoff	$2H_3O^+ + 2e^- \rightleftharpoons 2H_2O + H_2$	0V
Eisen(III)	$Fe^{3+} + 3e^- \rightleftharpoons Fe$	-0,037V
Blei	$Pb^{2+} + 2e^- \rightleftharpoons Pb$	-0,126V
Zinn	$Sn^{2+} + 2e^- \rightleftharpoons Sn$	-0,141V
Chrom	$Cr^{3+} + 3e^- \rightleftharpoons Cr$	-0,89V
Zink	$Zn^{2+} + 2e^- \rightleftharpoons Zn$	-0,762V
Aluminium	$Al^{3+} + 3e^- \rightleftharpoons Al$	-1,677V
Natrium	$Na^+ + e^- \rightleftharpoons Na$	-2,714V
Calcium	$Ca^{2+} + 2e^- \rightleftharpoons Ca$	-2,868V
Lithium	$Li^+ + e^- \rightleftharpoons Li$	-3,040V

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