

Die elektrochemische Spannungsreihe

| Element im Redoxpaar, dessen Oxidationsstufe sich ändert | oxidierte Form + $z\text{ e}^- \rightleftharpoons$ reduzierte Form | Standardpotential E^0 |
|--|---|-------------------------|
| Fluor | $\text{F}_2 + 2\text{e}^- \rightleftharpoons 2\text{F}^-$ | +2,89V |
| Chlor | $\text{Cl}_2 + 2\text{e}^- \rightleftharpoons 2\text{Cl}^-$ | +1,396V |
| Brom | $\text{Br}_2 + 2\text{e}^- \rightleftharpoons 2\text{Br}^-$ | +1,098V |
| Silber | $\text{Ag}^+ + \text{e}^- \rightleftharpoons \text{Ag}$ | +0,799V |
| Iod | $\text{I}_2 + 2\text{e}^- \rightleftharpoons 2\text{I}^-$ | +0,535V |
| Kupfer | $\text{Cu}^{2+} + 2\text{e}^- \rightleftharpoons \text{Cu}$ | +0,518V |
| Wasserstoff | $2\text{H}_3\text{O}^+ + 2\text{e}^- \rightleftharpoons 2\text{H}_2\text{O} + \text{H}_2$ | 0V |
| Blei | $\text{Pb}^{2+} + 2\text{e}^- \rightleftharpoons \text{Pb}$ | -0,126V |
| Zinn | $\text{Sn}^{2+} + 2\text{e}^- \rightleftharpoons \text{Sn}$ | -0,141V |
| Zink | $\text{Zn}^{2+} + 2\text{e}^- \rightleftharpoons \text{Zn}$ | -0,762V |
| Aluminium | $\text{Al}^{3+} + 3\text{e}^- \rightleftharpoons \text{Al}$ | -1,677V |
| Lithium | $\text{Li}^+ + \text{e}^- \rightleftharpoons \text{Li}$ | -2,040V |

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